

TUFTS UNIVERSITY
DEPARTMENT OF CHEMISTRY

The author wishes to extend his
sincere appreciation to Professor

PART I

Line broadening in the X-ray diffraction
patterns of the vanadium hydrogen system.

A Thesis
by
John H. MacMillan

LINE BROADENING IN X-RAY DIFFRACTION PATTERNS OF THE VANADIUM-HYDROGEN SYSTEM

- Introduction -

When Hydrogen was taken up by body-centered cubic Vanadium metal, an anisotropic expansion to a body-centered tetragonal lattice has been reported. (1)

(1) Arnulf J. Maeland - J. Phys. Chem. 68, 2197, (1964)

It is difficult to account for the observed expansion if hydrogen occupies only tetrahedral interstices and there is no convincing evidence that octahedral interstices are involved. (2)

(2) T.R.P. Gibb Jr. "Advances in Chemistry" 39, 99, (1963).

The problem could be resolved by neutron-diffraction but experimental observations by Hardecastle, (3)

(3) K. Hardecastle, unpublished results

and a study of the possibly comparable TaH_2 by Wallace (4)

(4) W.E. Wallace - J. Chem. Phys. 35, 6, 2158-2164 (1961)

have not led to useful results, partly because of the complexity of the diffraction patterns and partly because of

the assumption that octahedral sites are not occupied.

Objectives

It was desired to investigate the breadths of the tetragonally distorted cubic phase X-ray diffraction lines of various vanadium hydrides and to compare these with the corresponding lines of the pure metal. Then one could ascertain whether certain lines corresponding to certain planes in the crystal, were broadened more than others, and it was also hoped that from these observations some idea of the positions of the hydrogen in the tetragonally distorted body centered cubic lattice could be obtained. Also it was hoped initially that accurate quantitative measurements of the line broadening could be carried out.

A.J. Maeland (1) investigated the change in lattice parameters with hydrogen content in nonstoichiometric hydrides of Vanadium via X-ray diffraction using the powder method with a Straumanis² cylindrical camera. His collection of films from his study were left with Dr. T.R.P. Gibb of Tufts University and these films were used in my observations. Maeland observed that between stoichiometries $V-H_{.05}$ and $V-H_{.46}$ a b.c.c. cubic phase and a b.c.t. phase exist together and that at room temperature the cubic phase disappears at $V-H_{.46}$. At higher temperatures the cubic phase appears at stoichiometries up to $\sim V-H_{.9}$. Accordingly, on some films the breadths of body centered

cubic diffraction planes were measured and compared with the breadths of the pure metal, while on some high H content films at room temperature b.c.t. planes are compared to corresponding planes in the body centered cubic pure vanadium lattice.

Experimental Methods

Initially a photometer available in the laboratory was used to record the widths of the lines, but it was found that the machine gave variable results for the same line on different days and took much time to measure one film. Therefore this procedure was abandoned and recourse was made to visually measuring the breadths of the seven cubic phase lines for many different films. To accomplish this, a box with a white glass illuminated field was employed with a movable steel support for the film and a rider on the support with magnifying glass and measuring device.

Since it was necessary to measure the line breadths visually only an approximate value for the breadths was obtained. There can be errors of measurement and with some faint lines it is extremely difficult for the eye to judge where the line ends and the field begins. However, the majority of lines were measured to $\pm .01$ mm. Over the average of a large number of films, however, any trends should become apparent.

Sources of Line Broadening

Line broadening can be caused by either extremely

small particle size or stress or disorder in the metal. (5)

(5) N. F. Mott and H. Jones "The Theory of the Properties of Metals and Alloys" Clarendon Press, Oxford 1936.

Since particle size causes broadening only for particle diameters below 1000 Å and the particles in the V metal used are coarse according to Maeland's article (1), line broadening in these films should then be due to stress and disorder in the metal. Any mechanical stress in the metal should be annealed out upon heating to form the hydride, therefore any broadening observed in the hydrides diffraction lines must be due to stress and disorder caused by hydrogen atoms slightly distorting the lattice. A plane in the crystal greatly distorted by hydrogen atoms should be substantially broadened.

Experimental Results

The results of the measurements are tabulated in the figures. Each cubic phase line is numbered and the Miller indices of the plane it represents are given. It is seen that lines 3, 2 and 1 often were obscured due to scattered radiation darkening the front reflection region of the films. Therefore results from these lines are less reliable. It is seen that there is no general trend of increased broadening with increasing hydrogen content but that broadening does occur in most lines with respect to vanadium metal lines.

Also temperature differences have little effect on the breadth of the lines. Rough percentage broadenings of the lines relative to vanadium are given on the third data page.

From these data the most noticeable trend is that line 5, the 211 plane, is substantially more broadened than all the others. Lines four and one are less broadened than the others but these results are less certain due to many obscured lines. The single most consistent feature of the data is the greater broadening of the line corresponding to the 211 plane.

BREADTHS FOR VANADIUM HYDRIDES DIFFRACTION LINES COMPARED TO PURE VANADIUM. BREADTHS IN MM.

Films from A. Masland

0 = line obscured

F = Faint, V.F. = Very Faint.

Diffraction Plane:

Compound	T ° C	7	6	5	4	3	2	1
		110	200	211	220	310	222	321
V powder	Room	.04	.04	.06	.06	.09	.08 est.	.16
"	158	.04	.04	.05	.07	.07	0	.17
"	305	.04	.04	.05	.07	.07	0	.17
V-H.15	81°	.04	.06	.08	.08 ~	.13	0 ~	.23
"	135°	.04	.04	.07	.07 F	.08	0	.17
"	308	.04	.04	.06	0	.07	.05 F	.18
V-H.26	Room	.05	.05	.09	.07 F ~	.12	0	~.2
"	"	.07	.06	.09	~.08 F ~	.12	0	0
"	200° C	~.1	~.1	~.15	~.15 F	0	0	0
V-H.30	169°	.06	.07	.07	.06 ~	.1	0	~.2
"	311°	.04	.06	.06	.07	.08	0	.17
V-H.37	Room	.04	.04	.08	0	0	0	0
"	163°	.04	.04	.07	0 ~	.12	0	F but 7 V line
"	320°	.04	.04	.06	.067 best	.07	0	.17
V-H.43	Room	.05	.05	.08	0	~.1 F	0	~.22 F
"	144°	.06	.05	.07	0	0	0	> V line
"	300°	.05	.05 F	.06	.06 F	.08 F	0	.17
V-H.59	Room	.05	.05	.06	0	0	0	0
"	164°	.05	~.09 VF	~.1	0	F but 7 V	0	0
"	311°	.04	.04	.06	.07 ~	.09	.06 F	.17
V-H.69	152°	.06	.05 F	.07	0	7 V	0	> V
"	300°	.06	.07 F	.08	.07 F	~.1 F	0	~.18
V-H.76	Room	.05	~.07	~.1	0	0	0	0
"	167°	.05	~.1	~.12	0	> V	0	0

ROUGH PERCENTAGE BROADENING IN VANADIUM HYDRIDE CRYSTAL PLANES
DIFFRACTION LINES COMPARED TO PURE VANADIUM AT COMPARABLE TEMPERATURES.

The 222 plane is omitted since lines from it are nearly always obscured.

Diffraction Plans:

Compound	T, °C	⁷ 110	⁶ 200	⁵ 211	⁴ 220	³ 310	¹ 321
V-H, 15	81°	0%	50%	60%	15%	65%	11%
"	135°	0%	0%	40%	0%	11%	0%
"	308°	0%	0%	20%	--	0%	6%
V H, 26	Room	25%	25%	80%	0%	73%	25%
"	"	75%	50%	80%	11%	53%	--
"	200°	150%	175%	200%	110%	--	--
V H, 30	169°	50%	75%	16%	0%	57%	25%
"	311°	0%	50%	20%	0%	11%	0%
V H, 57	Room	0%	0%	60%	--	--	--
"	163°	0%	0%	40%	--	70%	10%
"	320°	0%	0%	20%	0%	0%	0%
V H, 43	Room	25%	25%	60%	--	43%	40%
"	114°	50%	25%	40%	--	--	10%
"	300°	25%	25%	20%	0%	11%	0%
V H, 59	Room	25%	25%	20%	--	--	--
"	164°	25%	--	100%	--	25%	--
"	311°	0%	0%	20%	0%	28%	0%
V, H, 69	152°	50%	25%	40%	--	25%	10%
"	300°	50%	75%	60%	0%	43%	5%
V-H, 76	Room	25%	75%	100%	--	--	--
"	167°	25%	150%	240%	--	25%	--
"	270°	50%	50%	60%	--	--	--
V-H, 80	Room	25%	25%	80%	43%	30%	--
"	Room	25%	50%	60%	11%	10%	--
"	200°	100%	50%	40%	--	--	--
"	158°	25%	50%	40%	0%	10%	--
Average		32%	43%	60%	11%	30%	12.5%

Conclusions

The greater broadening of the 211 plane relative to the others leads us to speculate about the positions of the hydrogen atoms in the lattice. In a body centered cubic lattice the hydrogen atoms can occupy either octahedral or tetrahedral sites. In a b.c.c. unit cell there are 24 tetrahedral and 18 octahedral sites.

(6) (See diagram one).

(6) T.R.P. Gibb Jr.-J.Phys.Chem. 68, 1096,(1964).

A great broadening would seem to indicate that hydrogens are in certain octahedral or tetrahedral sites. Let us now tabulate the octahedral and tetrahedral points for the various planes in one unit cell. (See diagrams).

<u>Plane</u>	<u>No. Octahedral</u>	<u>No. Tetrahedral</u>
211	11	0
110	4	0
222	3	0
321	1	0
310	1	0
200	8	8
220	6	4

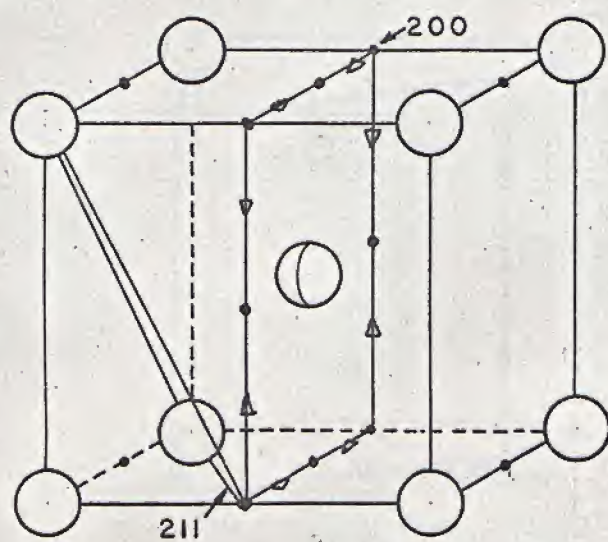
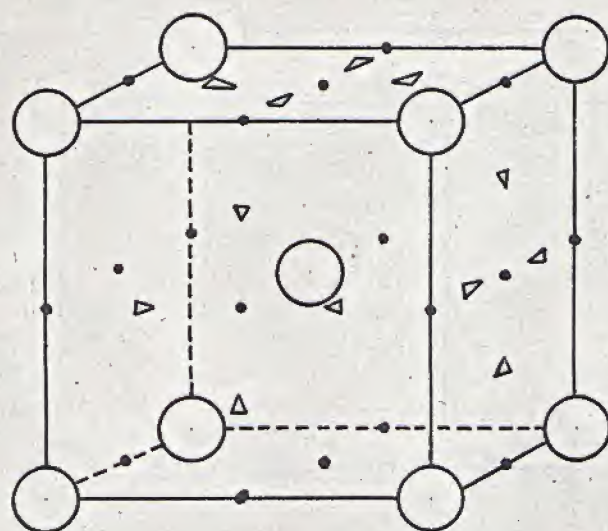
FIGURE IFirst Illustration:

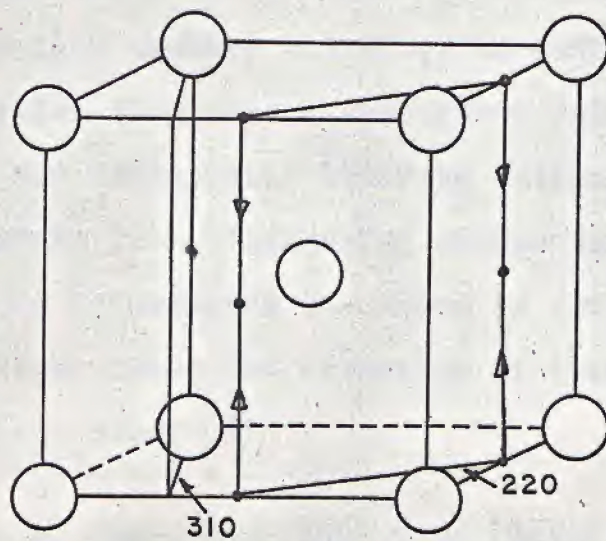
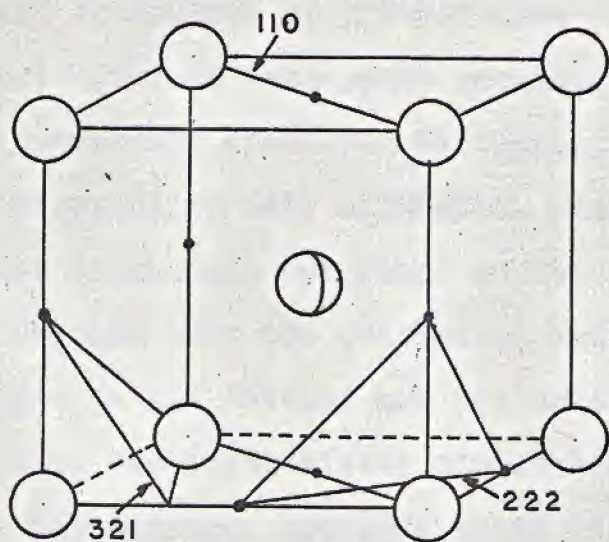
Octahedral (•) and tetrahedral (◐) points in b.c.c. unit cell. (Not all points are shown on every face).

Other Illustrations:

Diffraction planes in Vanadium b.c.c. lattice showing octahedral and tetrahedral intersections.







With just the percentage broadening information one cannot unequivocally distinguish between octahedral and tetrahedral occupation or between face-face or edge-edge site restrictions in octahedral occupation. The fact that so many appreciably broadened diffraction lines are from planes intersecting only octahedral points at first seems to indicate octahedral occupancy in the lattice. We must, however, realize that the octahedral and tetrahedral sites for hydrogen in the lattice are not mere mathematical points. Both the interstitial atomic and the hydridic model (3) for vanadium hydrides treat the hydrogen in the lattice as a small share of electron density. For the interstitial model the radius of the hydrogen sphere is $.55 \text{ \AA}$ for 4 coordination and $.60 \text{ \AA}$ for 6 coordination. In the hydridic model the H^- has a radius of 1.32 \AA (which increases with coordination number) therefore no matter which model is used, distortion of planes intersecting only octahedral points does not necessarily indicate octahedral occupation since a hydrogen in a neighboring tetrahedral site whose electron cloud intersected the plane in question could distort the lattice along the direction of that plane.

Roberts' N.M.R. study (7)

(7) Roberts, B. Phys. Rev. 100 1257 (1955)

of the V D₂ system indicates that the deuteriums are a $\frac{\sqrt{3}}{2}$ apart. This fact would support octahedral occupation if the hydrogen sites were mere geometric points, since no combination of octahedral and tetrahedral sites in a b.c.c. lattice is separated by that geometry. However with the hydrogen possessing a finite atomic radius, this simple argument breaks down. Such considerations make any meaningful correlation between the rough average broadening of the diffraction lines and octahedral or tetrahedral occupancy seem dubious. For example, the broadened 311 line could result from distortion of the plane by hydrogen occupying the octahedral site in that plane or by hydrogen from a neighboring tetrahedral site distorting the plane by overlap, since the hydrogen has a finite size.

Naively, we might expect the greatest percentage lattice deformation and greatest line broadening in the planes with the densest packing and smallest area in the unit cell. In the planes of greater area one could predict that the stress deformation of the V lattice along the plane would be less than for those of smaller area since the stress from the hydrogen would tend to fall off at greater distances from the hydrogen. Such a conclusion would be valid, however, only if the same number of hydrogens were acting to distort each plane, and only if interaction with electron clouds from neighboring sites were the same. Clearly this cannot

be assumed since some of the larger planes can accommodate two hydrogens on opposite faces or edges even when considering bulky H^- ions. That this effect is not applicable is also seen with the densely packed and small area 211, 310 and 321 planes all of which intersect only 1 octahedral point and no tetrahedral points. The 211 plane is evidently distorted more than the other two. It seems futile to speculate at present as to the cause for this since we know neither the configuration of hydrogen in the lattice nor the degree to which electron density from hydrogens adjacent to the planes may affect the distortion.

Introduction

The purpose of this work is to study the magnetic properties of a certain substance. The results of the experiments are given in the following sections.

(1) The effect of temperature on the magnetic properties of the substance.

Chicago, Illinois (1944)

PART II

Magnetic Susceptibility Effects on

Removal of Hydrogen from Phase

Palladium Hydride.

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METHODS

4. The author wishes to extend his
deep appreciation to Professor

14. T.R.P. Gibb Jr., without whose
patient guidance this paper

15. would not have been possible.

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X-RAY DIFFRACTION

Introduction

For decades it has been known that the magnetic susceptibility of Pd decreases as hydrogen is absorbed and reaches zero at approx. Pd H_{.6} to Pd H_{.7} (8).

(8) D.P. Smith "Hydrogen in Metals" University of Chicago Press, Chicago, Illinois (1948)

It has also been known for many years that at Pd H_{≤.03}. The Pd metal expands as a phase termed α that is only slightly greater in lattice parameter, and this phase co-exists with a second β phase of parameter approximately 4.02 Å up to stoichiometry \sim Pd H_{.6} whereupon the β phase appears solely. (8) Thus it was necessary to explain why the β phase was diamagnetic and the familiar band theory (5), was the most satisfactory at explaining the effect. According to this explanation the portion of the density of states curve for Pd to the right of the Fermi level shows that Pd lacks approximately .55 electron in its d band. When H₂ is absorbed the electron from the hydrogen is donated to the d band therefore at hydride composition of approximately Pd H_{.55} there should be no unpaired electron spin in the d band and the substance should be diamagnetic.

However Michel and Gallissot (9) removed all of the

(9) A. Michel and M. Gallissot: Compt. Rend. 208 434 (1939)

hydrogen from β phase palladium hydride with a high voltage device at room temperature and observed no increase in susceptibility. Thus they postulated that volume effects in the lattice, not filling of a hole in the d band was responsible for the decrease in magnetic susceptibility.

T.R.P. Gibb Jr. (10) noted that if one extrapolates the

T.R.P. Gibb Jr. "Lattice-Distension and the Magnetic Susceptibility of Palladium Hydride"- Tufts Univ. 1966, to be published

curve of paramagnetic susceptibility vs. relative volume of the unit cell the value at which pure palladium becomes diamagnetic is surprisingly close to the relative volume of the β phase lattice. T.R.P. Gibb and W.A. Norder attempted to confirm the results of Michel and Gallissot and their results were partially confirmed. Accordingly in this investigation Dr. R.J. Roy and myself wished to prepare pure β phase hydride and pump off hydrogen to some stoichiometry below $\text{Pd H}_{0.6}$ without forming any α phase and wished to observe if the magnetic susceptibility remained constant. A zero paramagnetic susceptibility at hydrogen contents lower than that of $\text{Pd}_{0.6}$ would refute the band theory explanation of this effect, and reinforce Michel and Gallissot and T. R. P. Gibb's speculations. If it were impossible

to prevent formation of ϵ phase we at least desired to obtain a graph of magnetic susceptibility vs. stoichiometry on desorption that would show a lower susceptibility than the corresponding stoichiometry on absorption. Such would infer that the desorption stoichiometry had a greater percentage of diamagnetic δ phase than the corresponding absorption stoichiometry.

Experimental

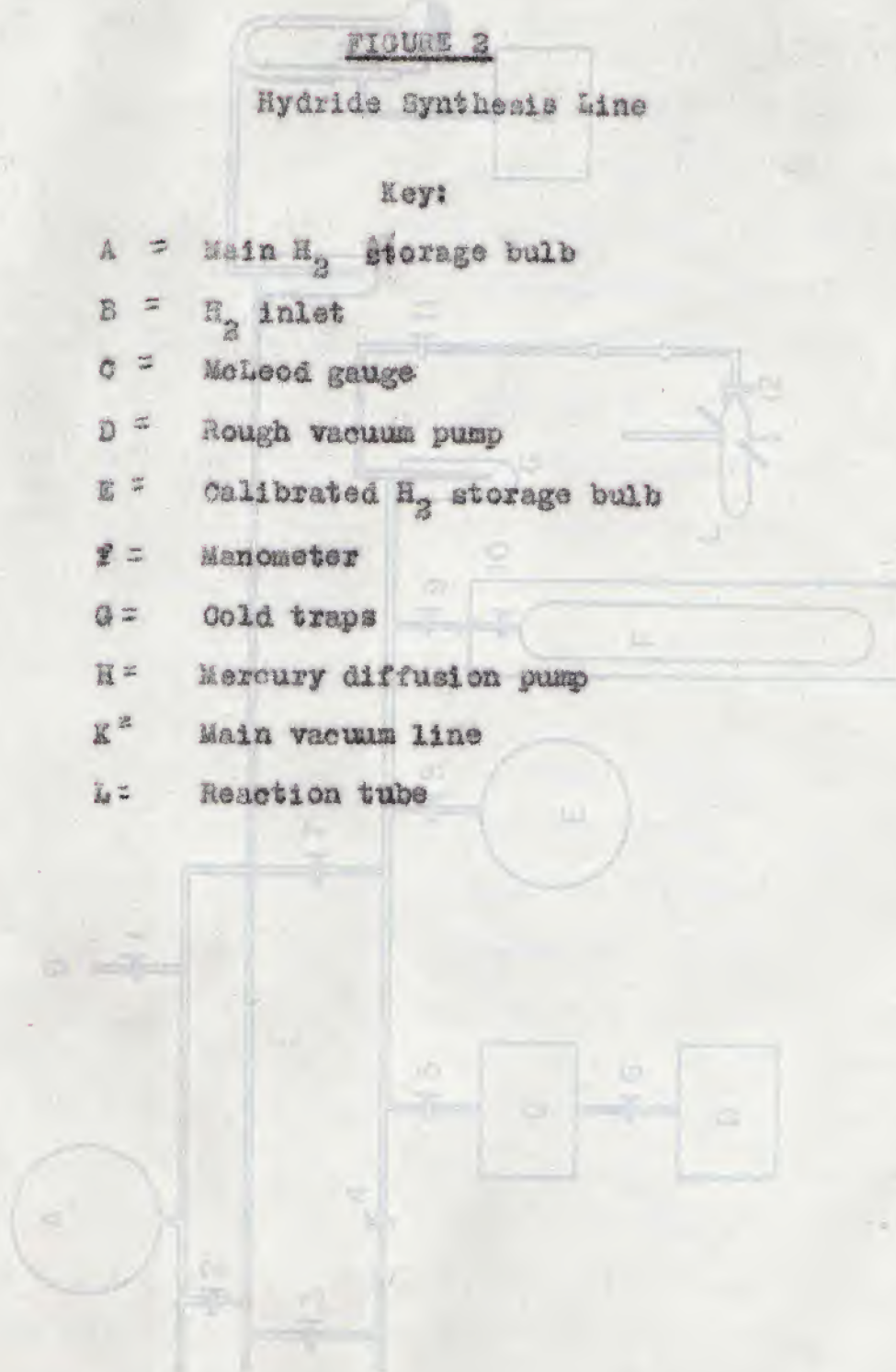
The hydrides were synthesized with the hydride synthesis line illustrated and described in the Figure 2. Tank hydrogen was passed through a catalytic Deoxo purifier and drierite before entering the system. The vacuum system could be evacuated to 10^{-5} mm. In the attempted earlier synthesis a $U H_3$ getter was used to generate the H_2 but it became very sluggish to charge and was finally abandoned. The early attempted syntheses employed an approximately 200 ml reaction tube and a metal boat for holding the palladium. When the synthesized hydrides were transferred in the inert atmosphere box to an analysis tube, the lock had to be evacuated with the hydride exposed. This procedure was found to remove nearly all of the hydrogen from the sample due to the hydrides appreciable equilibrium pressure in a vacuum at r. t. Subsequent analysis of these samples showed no H_2 evolved and X-ray powder patterns showed only

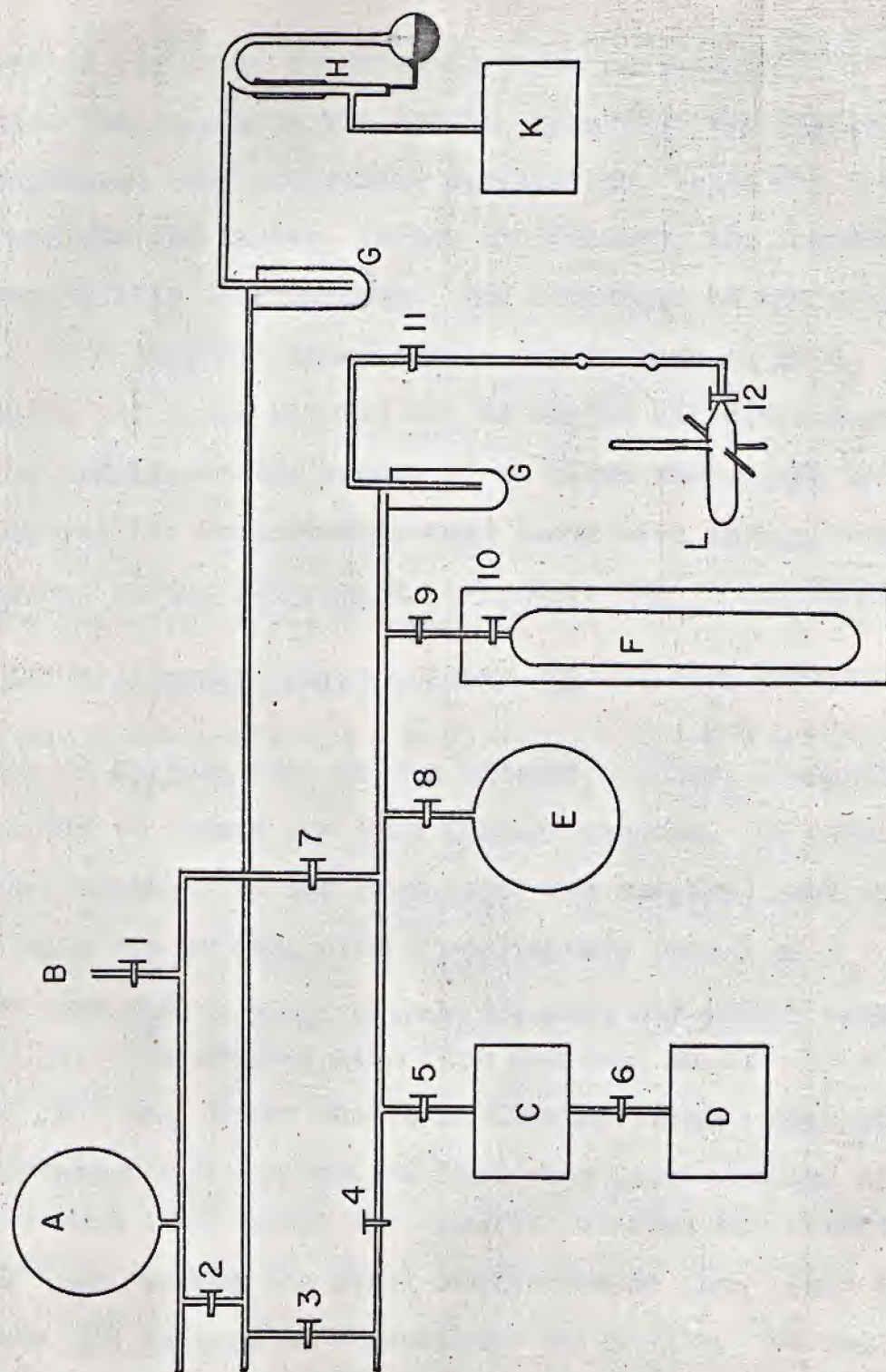
FIGURE 2

Hydride Synthesis Line

Key:

- A = Main H_2 storage bulb
 B = H_2 inlet
 C = McLeod gauge
 D = Rough vacuum pump
 E = Calibrated H_2 storage bulb
 F = Manometer
 G = Cold traps
 H = Mercury diffusion pump
 K = Main vacuum line
 L = Reaction tube





phase present. To overcome these difficulties the reaction tube shown in the hydride apparatus was designed. It contained four protruding capillaries, three for X-ray analyses and the fourth, larger in diameter, for magnetic susceptibility measurements. The advantage of the apparatus is that magnetic measurements may be made in an H_2 atmosphere and X-ray samples may be sealed off without exposing the hydride to the vacuum of an inert atmosphere box. Susceptibility measurements were taken with an apparatus described in the literature (11). Metz 99% .8 micron Pd used.

(11) W.A.Norder, Rev.Sci.Instr. 31 848-851 (1960).

The Pd filings were passed through a screen alongside a magnet to remove any iron filings present. In order to remove oxides from the surface of the samples, each weighed sample was treated with a preliminary volume of H_2 at r.t. to form H_2O at room temperature with any oxygen present as oxide. The samples were then degassed at 100° to a pressure of 10^{-5} mm. Every sample so treated formed noticeable amounts of H_2O and evolved heat when first exposed to dried hydrogen even though the reaction chamber was evacuated to 10^{-5} mm. before the first admittance of gas. Thus the palladium did contain an appreciable oxide film. It was found through trial and error with preliminary samples that Pd

absorbs negligible amounts of H_2 at r.t. but that absorption is appreciable at 60° to 80° C. Thus each weighed and previously decoxidized Pd sample was exposed to H_2 in the reaction tube from a calibrated volume of the vacuum line. The temperature was raised to 100° C. and the reaction tube allowed to cool at about 10° an hour to r.t. The stoichiometries were calculated from observing pressure changes as is described in a following section. Using this procedure it was possible to prepare hydrides of stoichiometries very close to the limiting ratios of Pd $H_{.7}$ at r.t.

An interesting observation is that on heating, appreciable sintering occurs when Pd metal is heated above 200° C. and that some sintering occurs even at 80° to 100° C. during hydriding of the Pd samples. Such behavior is quite surprising considering the high (over 1500° C) melting point of Pd. Due to this behavior the hydrides in the special reaction tube were in the form of a porous cake after the reactions. Since the mass couldn't be broken up manually without exposure to air, it was vibrated within the tube. A loudspeaker was used as the vibrator. It was modified by cementing a plastic hollow cylinder to the moving portion so that the plastic cylinder agitated the tube. An audio oscillator and amplifier provided vibration of the proper frequency so that the porous mass within the tube was reduced to powder after about two days of vibration. Such procedure was necessary in order for the hydride to be transferable into the four capillaries of the

reaction type.

Another interesting observation is the intensity of the exothermic reaction when the hydrides are exposed to air. Water is immediately formed and so much heat is evolved that small particles glow in the reaction tube. The reaction could be due either to oxygen reacting directly with the hydride, or to metallic Pd catalyzing the reaction between O_2 and H_2 evolved from the surface of the hydride.

(12) Meyer, H. Z. Elektrolyse, 1927.

$$V = \frac{P_1 V_1}{P_2} - (V_2 - V_1)$$

V = volume of gas

V_1 = volume of gas before expansion

V_2 = volume of gas after expansion

P_1 = pressure of gas before expansion

P_2 = pressure of gas after expansion

$V_1 = V_2 + V_3$ (1-20) (1-20)

V_3 = volume of gas in the tube

Calculated values

Table 1 + 2 = 1 + 2 + 12 (see Figure 1)

0.017 g

Volume of reaction tube = 10.0 ml.

CALIBRATION OF THE SYSTEM

The volume of two sections of the vacuum system was obtained by expanding a known pressure and calibrated volume, known calibrated bulb $V_c = 938.9$ ml.

The following formula was used: (13)

(13) Hayes, H. M.S.Thesis, Tufts University, 1957

$$V_1 = \frac{P_c V_c}{P_1} - (V_c - h_1)$$

V_1 = unknown volume

P_c = Pressure in calibrated flask before expansion

P_1 = Pressure in calibrated flask + unknown volume after expansion

V_c = Volume of calibrated flask

h_1 = correction for volume taken up by the Hg.

$$h_1 = \pi r^2 (h-50) = .3827 (h-50)$$

h being height in cm of the Hg.

Calibrated volumes

Volume 4 + 5 + 7 + 8 + 12 (see Figure 2)

= 923.7 ml

Volume of reaction tube L = 49.4 ml.

CALCULATION OF HYDRIDE STOICHIOMETRY

The stoichiometry of the hydride may be computed during the synthesis by observing the decrease in pressure as the known volume of hydrogen is admitted to the reaction chamber. All pressure readings must of course be made at the same temperature (room temperature in these runs).

$$P_1 V_1 = (V_2 + V_3 + V_4) P_2$$

P_1 and V_1 = pressure and volume in calibrated system.

P_2 = pressure in system + reaction tube after reaction

$V_2 + V_3 + V_4$ = total effective volume after reaction

V_2 = known volume reaction tube

V_3 = volume H_2 absorbed by Pd at P_2 pressure and 298° K temperature.

$$\therefore \frac{P_2 V_3}{298} = 76 \times \frac{V_4}{273}$$

V_4 = Volume absorbed at S.T.P.

$$E_1 = \frac{V_4 \text{ ml}}{1.1207 \times 10^4 \text{ ml/equiv.}} = \text{equivalents } H_2 \text{ absorbed}$$

$$E_2 = \frac{\text{wt Pd}}{106.7} = \text{equivalents Pd}$$

$$\text{Stoichiometry} = Pd H_{\frac{E_1}{E_2}}$$

RESULTS

In order to take the hydrogen out of the sample under mild enough conditions to allow some survival of the β phase on desorption, the sample was exposed repeatedly to a calibrated line under vacuum next to the magnetic susceptibility apparatus (V of line = 1718 ml). By reading with a manometer the pressure on the line when the evacuated system is exposed to the reaction tube, we can calculate the volume released by the sample and from this the stoichiometry of the hydride in the reaction tube. After the number of expansions came to a total of at least 1 cm pressure the hydrides were placed in the magnetic susceptibility side-arm of the reaction tube and the susceptibility measured.

Due to experimental difficulties, accidents, and a time limit on this thesis there was time for only one run with a diamagnetic phytide sample. One early sample showed appreciable paramagnetism (1.714×10^{-6} g.s/g) although the stoichiometry came to Pd H_{0.68} and therefore it was discarded. The stopcock popped on another promising sample thus greatly delaying compilation of data.

In the sample that was finally studied 2.3528 g of Pd absorbed 171.9 ml of H₂ at S.T.P. corresponding to a stoichiometry of Pd H_{0.696}. The sample was subjected to vibration to break down the porous mass. Susceptibility measurements

showed the substance to be diamagnetic as is predicted in all the literature on magnetic effects of the PdH_2 system(5). The results of the data are shown in the forthcoming table. Correction is made in the first reading for the pressure of H_2 initially over the Pd H.696 (28.5 ml at S.T.P.) The pressures obtained upon expansion into the calibrated system were quite high at first but after 30 expansions the pressure of the H_2 became negligible. The total calculated volume given off by the sample (203ml) agreed quite well with the known total volume of H_2 at S.T.P. (200.4ml), considering the limited accuracy of reading the manometer.

CONCLUSIONS

The most striking part of the data is the result that the hydride remains diamagnetic at least until stoichiometry Pd H.₃. This effect strongly indicates that under the mild conditions of evacuation employed here the diamagnetic phase is slow to break down to paramagnetic and therefore at stoichiometries at least to Pd H.₃ the diamagnetic phase exists solely. Unfortunately since this run was originally meant to be a trial and I did not know what behavior would be observed, X-ray pictures were not taken during the run to confirm this. It is hard to explain diamagnetism for stoichiometries <.6 in any other way though since the phase, which is only slightly distended Pd metal, is

MAGNETIC SUSCEPTIBILITY VS. STOICHIOMETRY FOR DEGASSING PdH₆₉₆
INTO A CALIBRATED SYSTEM. V of H₂ over PdH₆₉₆ initially=28.5 ml at S.T.P.

Time between evac.	Evacua- tion	P (cm)	Vol. H ₂ S.T.P.	Equivalents H ₂ removed	Total Equivalents Removed	M/Pd	X Mx10 ⁻⁶	X N Pd
	1	2.02	41.84	1.19x10 ⁻³	1.19x10 ⁻³	.61	0	0
30 min.	2	2.30	47.84	4.25x10 ⁻³	5.44x10 ⁻³	.45	0	0
30 min.	3	.80	16.57	1.48x10 ⁻³	6.92x10 ⁻³	.38	0	0
Overnight	4a	.37						
10 min.	4b	.37						
10 min.	4c	.36						
20 min.	4 total	1.10	22.79	2.04x10 ⁻³	8.96x10 ⁻³	.29	0	0
Overnight	5a	.36						
10 min.	5b	.36						
10 min.	5c	.36						
20 min.	5 total	1.08	22.34	1.99x10 ⁻³	1.095x10 ⁻²	.2	190.5	.354
Overnight	6a	.35						
10 min.	6b	.35						
10 min.	6c	.35						
20 min.	6 total	1.05	21.74	1.94x10 ⁻³	1.289x10 ⁻²	.11	204.8	.381
30 min.	7a	.16						
10 min.	7b	.16						
10 min.	7c	.30						
10 min.	7d	.21						
10 min.	7e	.21						
10 min.	7f	.20						
80 min.	7 total	1.20	25.89	2.31x10 ⁻³	1.520x10 ⁻²	.01	217.2	.404
Overnight	8a	.17						
10 min.	8b	.10						
10 min.	8c	.13						
10 min.	8d	.04						
10 min.	8e	.01						
40 min.	8 total	.45	10.8	8x10 ⁻⁴	1.61x10 ⁻²	.001	277.2	.515
Overnight	9a	.000	0	0	0	.000	321.2	.558

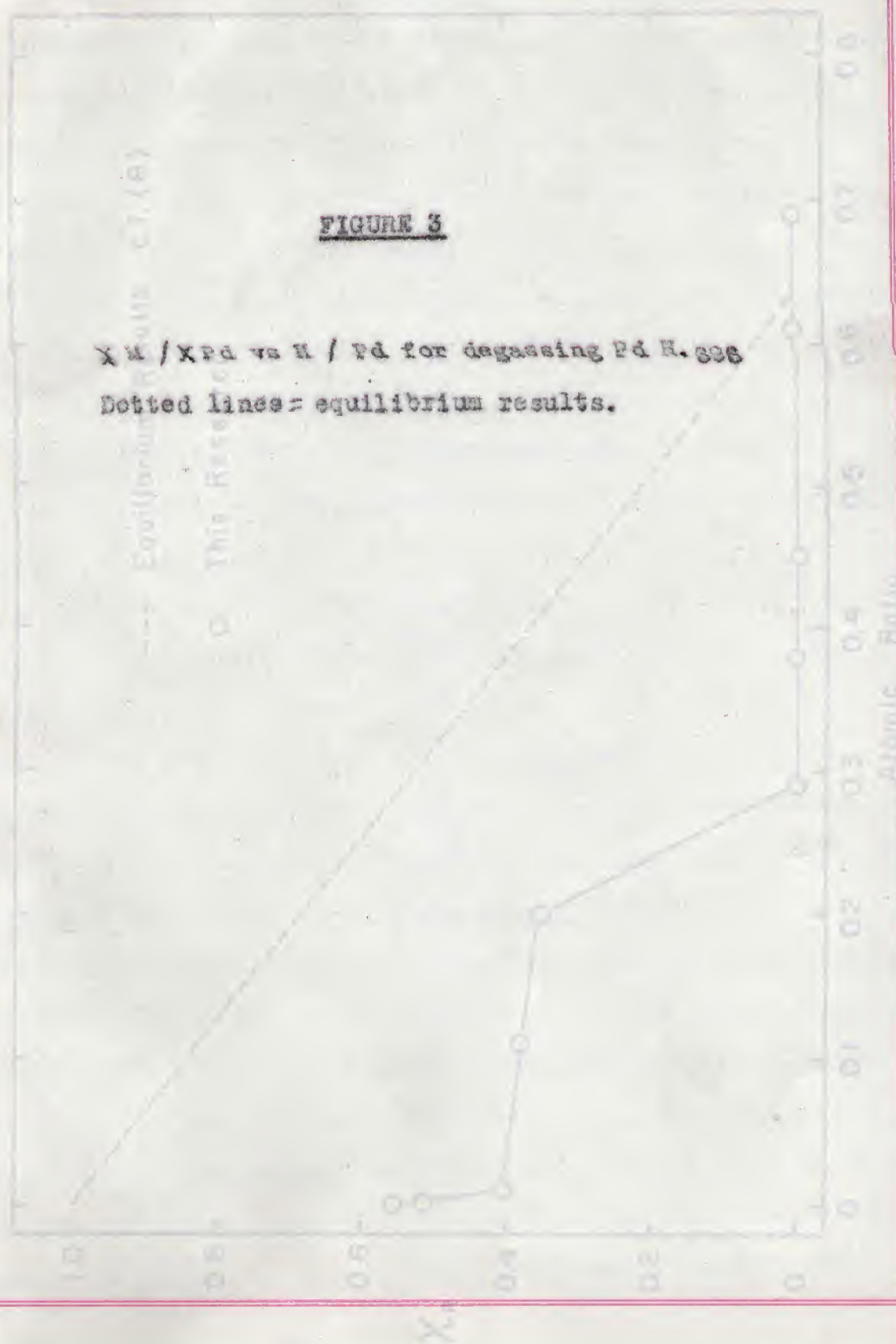
unquestionably paramagnetic. Pumping off substantial quantities of H_2 without changing the susceptibility confirms the work of Michel and Gallissot (9). This observation also tends to support lattice distension of Pd as proposed by Gibb (10) as the reason for the drop in susceptibility as H_2 is absorbed. It is not consistent with the proposal that filling of the approximately .6 hole in the d band causes the drop in paramagnetism, because were this the case pumping hydrogen off from stoichiometry .6 on down would remove spin pairing hydrogen electrons from the band causing an appreciable paramagnetic moment in the range (.6 to .3) where the hydride was found to be diamagnetic.

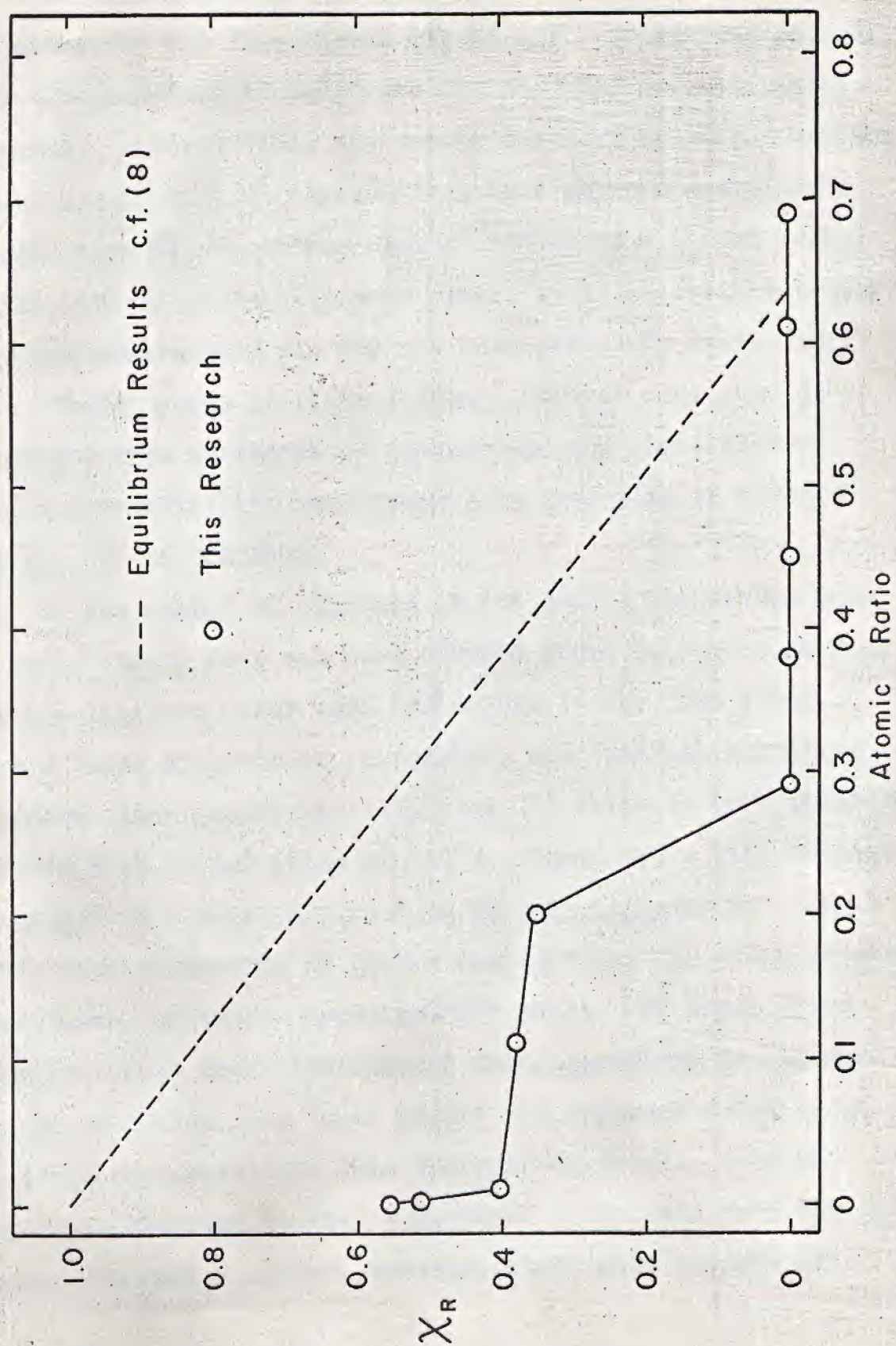
The accompanying graph illustrates these observations pictorially. The large increase in magnetic susceptibility on desorption from $Pd H_{.29}$ to $Pd H_{.2}$ indicates that in this range considerable defect β phase has broken down to α phase. ~~$Pd H_{.29}$~~ \rightarrow α two-phase system probably exists between $Pd H_{.29}$ and $Pd H_{.01}$. The relatively slight increases in magnetic susceptibility in this region indicate that comparatively small amounts of β phase are breaking down to α phase. It is not known whether the relatively great increase in susceptibility followed by smaller increases is a result of the experimental conditions or due to a fundamental phase change in the lattice in this stoichiometric region. All samples were treated under the same temperature (room) and

FIGURE 3

X_H / X_{Pd} vs H / Pd for degassing Pd H. 898

Dotted lines: equilibrium results.





approximately the same vacuum ($\approx .05$ mm). Since long periods were often required to calibrate the magnetic susceptibility apparatus, several times the sample was left overnight between evacuations. Such an interval may have allowed more β phase to decompose ~~to α~~ in the region between Pd H._{2.29} and Pd H.₂ if the rate of conversion were slow. It is impossible to say from one preliminary run why the susceptibility varies as it does after α phase is first formed. Further runs should be conducted with variances in temperature and time between taking susceptibility measurements to ascertain if these factors are significant.

As the amount of hydrogen in the sample approaches zero we would expect more and more defect β phase to decompose to α until a limiting value of α to β exists in the pure metal. Such a value probably is temperature and reaction condition dependent since Michel and Gallissot (9) claim to have obtained pure Pd with 100% β phase at -40°C . They used a high voltage discharge to remove the hydrogen at this temperature. It is known that desorption at higher temperatures and other drastic conditions produces appreciable α phase. If any β phase survived under the conditions of this experiment in the degassed palladium we would expect the degassed metal to be of lower susceptibility than the pure Pd metal. Such an effect is observed in this experiment as is seen from the graph. We must remember, however, that small amounts of

unevacuated H_2 remaining in the Pd also could lower the susceptibility somewhat. X-ray analysis of the degassed sample showed strong α phase lines and three very faint lines in the front reflection region that at first were thought to be β phase. Accurate calculations however gave $a_0 \approx 4.3 \text{ \AA}$ not ≈ 4.03 therefore the nature of this phase is unknown.

Since no β phase was detected and since the unknown impurity is evidently present in very small quantity the low susceptibility can probably be attributed to residual hydrogen. This conclusion is by no means certain however. The low hydrogen stoichiometry region of the graph is somewhat inaccurate due to the low hydrogen pressures and the difficulty of reading them accurately. This inaccuracy in no way detracts, however, from the more accurately determined stoichiometries from Pd $H_{0.69}$ to Pd $H_{1.1}$. The results in this region support the dependence of magnetic susceptibility on the Pd-Pd distance, and casts doubt on the Mott and Jones explanation for the magnetic behavior (5). Since only one run was made, one must regard these results as tentative. More runs should be undertaken to ascertain at what stoichiometry the susceptibility increases from zero as hydrogen is removed from Pd $H_{0.6}$. The effect of varying temperature and time between desorptions, and what effect very small amounts of hydrogen have on the susceptibility of palladium metal, also should be investigated.

Respectfully submitted,

John H. Macmillan

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CHAIRMAN,
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The Magnetic Susceptibility of Palladium Hydride

Sir: The linear decrease of the paramagnetic susceptibility of palladium as it absorbs hydrogen is well known.¹ The portion of the density of states curve to the right of the Fermi level is usually drawn to indicate that palladium lacks about 0.55 electron to fill the d level. It has been suggested² that hydrogen donates its electron to fill this d level so that when the spins are paired, the substance should become diamagnetic. This latter is observed experimentally but at a composition variously reported but close to $\text{PdH}_{0.66}$ rather than at $\text{PdH}_{0.55}$. If the above explanation is correct, it should be possible to obtain a linear increase in susceptibility as hydrogen is removed from $\text{PdH}_{0.66}$. This was attempted³ by a high-voltage method for extracting hydrogen without heating the sample and it was claimed that all the hydrogen could be removed without any change in the magnetic susceptibility. These results were not confirmed by Lewis, *et al.*,⁴ using an electrolytic method for removing hydrogen but were partially confirmed in this laboratory⁵ using the high-voltage method.

In the present series of experiments diamagnetic and slightly paramagnetic samples of palladium hydride were prepared from very fine (0.80- μ) palladium metal powder, by alternately heating and cooling the metal in pure hydrogen (obtained by evolution from UH_2), at a temperature which never exceeded 200° above which the powder sinters. These samples evolved hydrogen at room temperature when the hydrogen pressure above them fell below 18 mm. The magnetic susceptibility of samples of palladium hydride was continuously compared to that of a standard (Mohr's salt), while small measured quantities of hydrogen were removed from the sample, in an apparatus which has been described previously.⁶ The results are shown in Figure 1. Curves 1 and 2 show that it is possible to remove a large fraction of the hydrogen from diamagnetic palladium hydride and from palladium hydride which is slightly paramagnetic owing to an initial lower

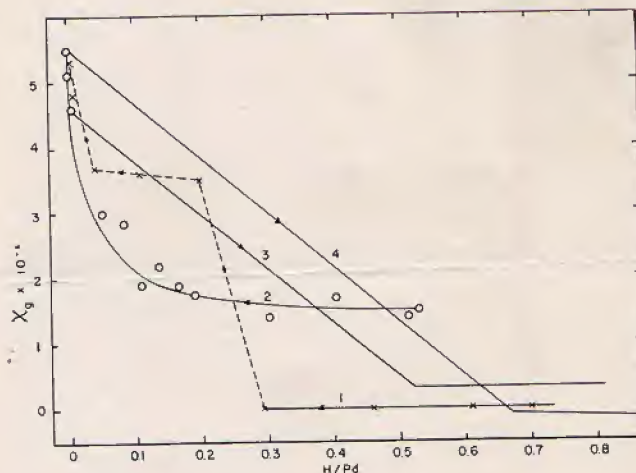


Figure 1. The effect of absorption and desorption of hydrogen on the paramagnetic susceptibility of palladium.

hydrogen content, without changing the susceptibility (χ_g). Curves 3 and 4 for the absorption of hydrogen by palladium are taken from Smith.¹ It thus appears that the band theory explanation is not tenable for the desorption of hydrogen from palladium hydride and it may be that lattice expansion plays a more important role than was formerly thought. Further experimental and theoretical work is in progress.

Acknowledgment. The authors are indebted to the U. S. Atomic Energy Commission for financial support for this work.

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